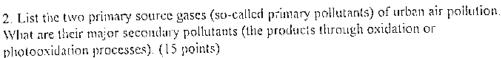
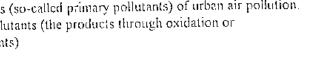
国立中央大學八十三學年度研究所碩士班入學試題卷

麦氧物理研究所 系所別:

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- 1. This problem is about the atmosphere.
- (a) What are the three major constituents of the atmosphere and what is their relative proportion? Roughly calculate the molecular weight of the air. (5 points)
- (b) Near the Earth's surface, if the atmospheric condition is at the temperature of 15 °C and pressure of 1 atm, what is the number density of the air in the unit of molecule cm-3? Given the universal gas constant = 0.082057 L-atm mole⁻¹ k⁻¹. The Avogadro's number is 6.022×10^{23} . (10 points)





- 3. A 40.0 ml sample of 0.1 M acetic acid (CH₃COOH) solution was titrated with 0.1 MNaOH. Calculate the pH of the solution under the following conditions.
- (a) After the addition of 10.0 ml of NaOH (10 points)
- (b) After the addition of 40.0 ml of NaOH (10 points)

Given $K_a = 1.754 \times 10^{-5}$ for CH₃COOH.

4. Given the following balanced equation

$$A \leftrightarrow B + C$$

and equilibrium constant $K = 5x \cdot 0^{-6} M$ at 25 °C.

What is the concentration of C at equilibrium at 25 $^{\circ}$ C when [A] = 5 M at the beginning of the reaction? Make any necessary assumption. (20 points)

5. A unimolecular decomposition has the following form,

$$A \longrightarrow B + C$$
 reaction constant = k.

Given the $m = [\Lambda]_0$, where the subscript "o" denotes the initial concentration of species Λ at t = 0.

- (a) Derive a general form of solution for [A] at reaction time t. (10 points)
- (b) For a given reaction the half-life t_{H2} is defined as the time required for its concentration to reach a value that is half-way between its initial and final values. Calculate the $|t_{I/2}|$ for the above reaction, (10 points) $|\cdot|$
- (c) The thermal decomposition of peroxy-acetyl nitrate (commonly called "PAN") is in a way as described above. If the reaction constant of PAN is 3,6x10-4 sec-1 at 25 °C, what is the chemical lifetime of PAN? What is the half-life of PAN? (10 points)

